

# EDTA - A Molecule with a Complex Story

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Molecule of the Month - March 2004

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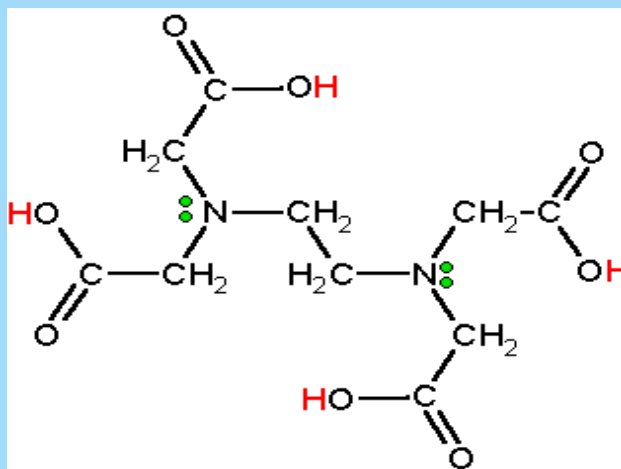
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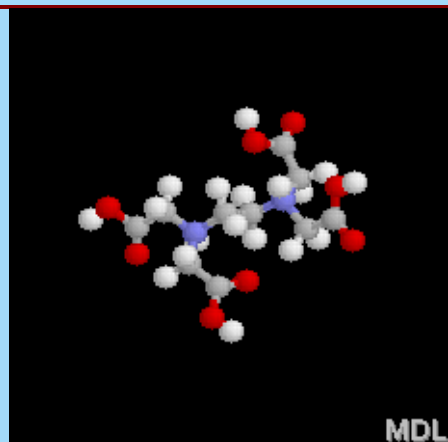
What do beer, the special sauce on a McDonald's Big Mac, and blue colored shampoos such as Aloe Vera or Revlon Aquamarine have in common?

## The EDTA Molecule

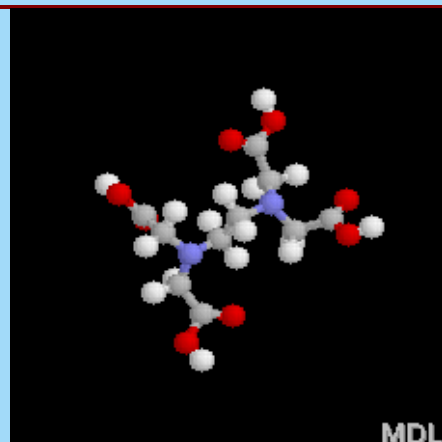
EDTA or ethylenediaminetetraacetic acid is a novel molecule for complexing metal ions. It is a polyprotic acid containing four carboxylic acid groups (acidic hydrogens are red) and two amine groups with lone pair electrons (green dots). The classic structural formula is given below. EDTA is synthesized on an industrial scale from ethylenediamine, [formaldehyde](#), and a source of cyanide (HCN or NaCN). [Click here](#) for the industrial reactions to open in a new window. On a worldwide basis over 100,000 metric tons are produced annually.



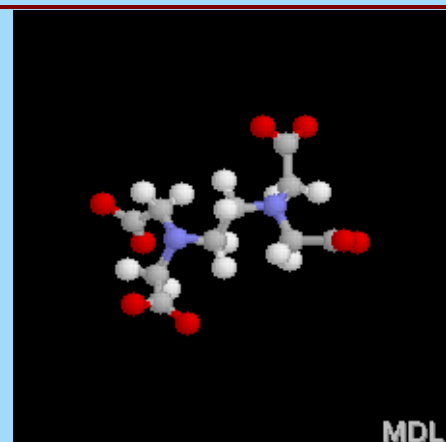
Besides the four carboxylic group hydrogens, EDTA can add two more hydrogens onto the amine groups. The structures of the fully protonated form (left), the typical form found in many textbook (center, matching the 2D structure above), and the fully deprotonated (all acidic H's removed) form (right) are given below.



forms at very low pH or  
very acidic condition  
(fully protonated)  
 $H_6Y^{+2}$



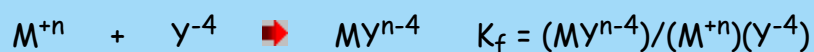
classic form  
 $H_4Y$



forms at very high pH or  
alkaline conditions  
(fully deprotonated)  
 $Y^{-4}$

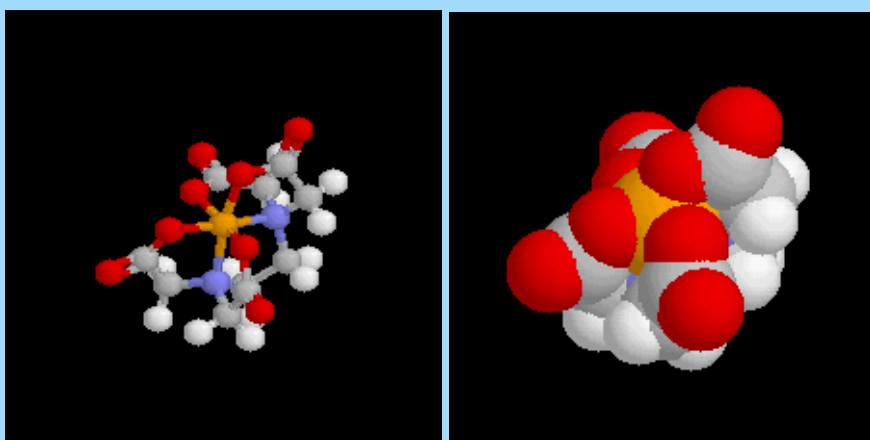
## Complexing Metals

The unusual property of EDTA is its ability to chelate or complex metal ions in 1:1 metal-to-EDTA complexes. The fully deprotonated form (all acidic hydrogens removed) of EDTA binds to the metal ion. The equilibrium or formation constants for most metals, especially the transition metals, are very large, hence the reactions are shifted to the complex. Many of the reactions are pH dependent, especially the weaker forming complexes with  $\text{Ca}^{+2}$  or  $\text{Mg}^{+2}$ .

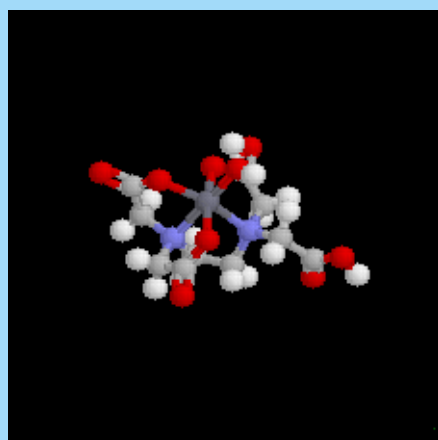


Metal analysis can be done by titration with EDTA and the use of a metal ion indicator. The pioneering work with EDTA was done by Gerold Schwarzenbach in the 1940's. The common reagent for making EDTA solutions is  $\text{Na}_2\text{H}_2\text{Y} \cdot 2\text{H}_2\text{O}$ . [Click here](#) for a list of formation constant ( $K_f$ ) values. The values of  $K_f$  increase with the charge on the metal ion and as ionic radius decreases with constant charge. Water hardness, mostly from dissolved  $\text{Ca}^{+2}$  and  $\text{Mg}^{+2}$ , is determined by EDTA titration at pH = 10 - [click here](#) for a typical experiment.

The structure of a classical complex of  $\text{Fe}^{+3}$  with EDTA is shown below. This is EDTA acting as a hexadentate ligand or all six sites on the EDTA bind to the metal ion. How would you describe the geometry of the octahedral  $\text{Fe}^{+3}$  in the complex?

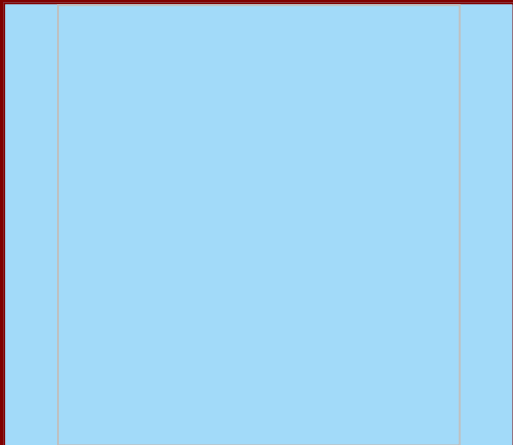
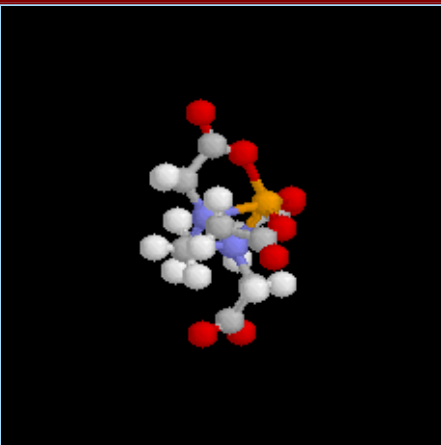
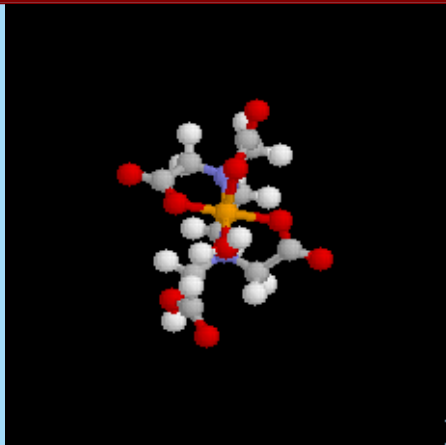


The octahedral coordination of the Fe-EDTA (left) and many other such complexes are very strained. The space-fill version of the structure (right) shows how crowded the structure appears. Consider the structure of  $\text{Cr}^{+3}$  with EDTA shown below. What is different for the Cr-EDTA complex compared to the Fe-EDTA complex?



A number of metal-EDTA complexes have been reported to have the EDTA acting as a pentadentate ligand (only five sites on EDTA bind, one carboxylic group does NOT). A water molecule or another ligand is in the sixth site, so the complexes are still octahedral in geometry.

So let's examine the formation of the complex. The animation below was produced in Spartan '04 by placing an  $\text{Fe}^{+3}$  (green) next to the EDTA with acidic H's removed (fully deprotonated as  $\text{Y}^{-4}$ ) and then minimizing the energy. Note that EDTA forms only five bonds to the  $\text{Fe}^{+3}$ , as the complex forms (center image is complex at end of animation). The image to the right has a molecule of water added, plus the EDTA has picked up a hydrogen ion on the carboxylic acid group not bound to the iron, and energy minimized to form the six coordinate complex.

		
<p>Animation of <math>\text{Y}^{-4}</math> complexing <math>\text{Fe}^{+3}</math>  <small>This is an avi file, so you'll need to have Real Player or Windows Media Player configured accordingly in order to see it.</small></p>	<p>complex with EDTA as a pentadentate ligand</p>	<p>octehedral complex with EDTA and water</p>

## Uses of EDTA

Because of its strong complexing ability for most metal ions, it is used in the food industry as a sequestering agent. The complexing of the metal ion may prevent further reactions, such as binding metals that are cofactors for enzymes, or just remove a metallic taste, such as metal contamination added during processing. See the Dow Chemical site on the use of their commercial product - [Versene](#). Some typical examples are given below.

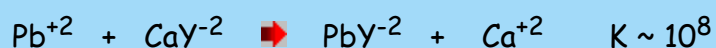
[Big Mac Special sauce](#)

[Canned beans](#)

[Hellmann's Mayonnaise](#)

Recent studies have shown that  $\text{NaFeEDTA}$  and  $\text{Na}_2\text{EDTA}$  added to typical iron fortification compounds in cereals increased the absorption of iron in adult humans.

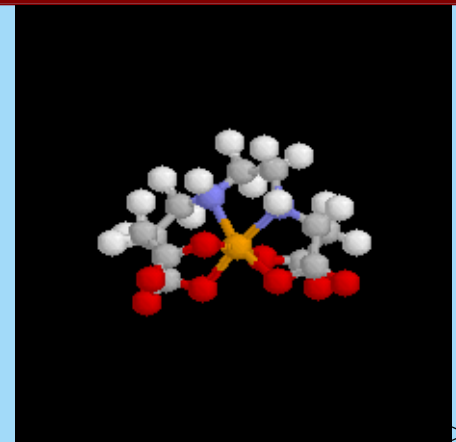
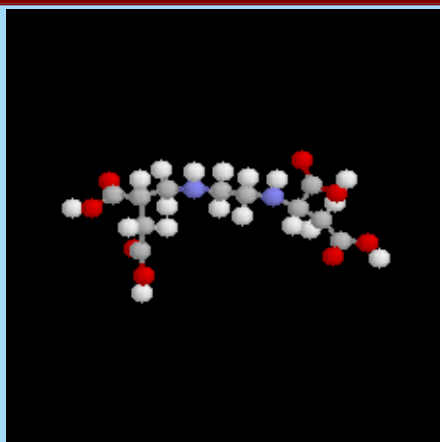
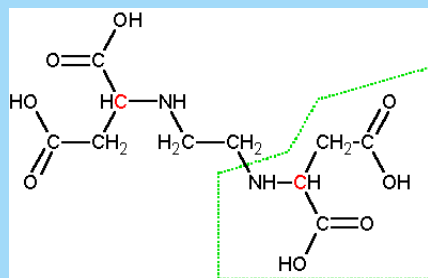
This same property allows EDTA use for incidents of lead poisoning by the medical profession. The formation constant for  $\text{Pb-EDTA}$  complex is  $10^{18}$ . Intravenous injection of  $\text{Na}_2\text{CaEDTA}$  solution is given at 25 mg/kg body mass/day over 6 hours for 5 days when blood lead levels go over  $45 \mu\text{g/dL}$ . The  $\text{Pb}^{+2}$  ion replaces the  $\text{Ca}^{+2}$  ion in the complex because the formation constant for the lead complex is greater than the calcium complex.



The five day limit is there to prevent  $\text{Zn}^{+2}$  depletion, since the  $\text{Zn}^{+2}$  ion replaces the  $\text{Ca}^{+2}$  ion in the complex too. EDTA is added to stored blood in blood banks as an anticoagulant to bind  $\text{Ca}^{+2}$  ion. Other reported uses of EDTA in medicine do not have a proven clinical basis. [Click here](#) for more information.

Another major use of EDTA has been in detergents to act as a builder (chelates metals) especially as a replacement for phosphates, a major nutrient in wastewater. However, a problem with EDTA is its inability to biodegrade in the environment. EDTA is found in many natural waters and occurs at higher levels in wastewater effluents. Western European countries have banned the use of EDTA in detergents. For a complete discussion of the chelating agents in the environment - [click here](#). EDDS ( $\text{S}$ ,  $\text{S}'$  -

ethylenediaminedisuccinic acid), a structural isomer of EDTA, has been used as a biodegradable substitute. EDDS is a good complexing agent and is broken down during wastewater treatment processes. This is a contribution to green chemistry by making a minor structural change. The cumulative production of EDDS by [Innospec Inc.](#) (formerly Octel) surpassed 10,000 metric tons in 2002.



EDDS structure showing chiral carbons in **red** - both must be *S*. *L*-Aspartic acid in **green** box.

biodegradable *S, S'* - EDDS molecule

octahedral FeEDDS<sup>-</sup> complex

EDTA is added to many commercial beers to stabilize foaming, taking advantage of the surfactant properties of EDTA, and used to remove scale by complexing calcium from calcium carbonate that forms on the processing equipment.

The blue color of the Cu-EDTA complex is used in many shampoos. See the [FDA use](#) of disodium EDTA-copper in cosmetics.

## References

Ethylenediaminetetraacetic Acid and Related Chelating Agents in *Ullmann's Encyclopedia of Industrial Chemistry* Vol. A10

What's that stuff: Food Preservatives, *C&EN*, 11 November 2002

<http://pubs.acs.org/cen/science/8045/8045sci2.html>

CSPI's Guide to Food Additives

<http://www.cspinet.org/reports/chemcuisine.htm>

The Brewer's Handbook

<http://www.beer-brewing.com>

Lead Poisoning in *The Merck Manual*

<http://www.merck.com/mmpe/sec21/ch326/ch326j.html>

UK Awards for Green Chemical Technology to Octel Performance Chemicals

[http://www.chemsoc.org/pdf/gcn/Octel\\_Award.pdf](http://www.chemsoc.org/pdf/gcn/Octel_Award.pdf)

A Brief History of Inorganic Classical Analysis

<http://www.cstl.nist.gov/nist839/839.01/beckhist.pdf>

Storing Up Trouble

<http://www.chemsoc.org/chembytes/ezine/1998/store.htm>

EDTA: The Chelating Agent under Environmental Scrutiny

<http://www.scielo.br/pdf/qn/v26n6/a20v26n6.pdf>

Chemical Speciation of EDDS and its metal complexes

<http://www.scilet.com/Papers/csb/csb113/CBSWhitburn.pdf>

An Evaluation of EDTA Compounds for Iron Fortification of Cereal-based Foods

*Br. J. Nutr.* 84 (6), 903-910 (2000)

Emerging Chemical Drinking Water Contaminants in *Identifying Future Drinking Water Contaminants*

<http://www.nap.edu/books/0309064325/html/112.html>

*Biogeochemistry of Chelating Agents*, edited by B. Nowack and J. VanBriesen, ACS Symposium Series Vol.

910, 2005 [http://www.ito.ethz.ch/SoilProt/staff/nowack/book\\_toc.html](http://www.ito.ethz.ch/SoilProt/staff/nowack/book_toc.html)

Just for FUN: See the EDTA song - [click here](#).



Feedback welcome - click on mail box.

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